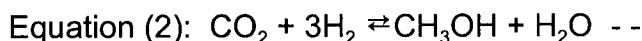


**Amendments to the Specification:**

Please replace the paragraph beginning at page 7, lines 24, with the following rewritten paragraph:

- - The term "equilibrium limited chemical reaction" refers to a chemical reaction or a set of complementary reactions that do not proceed to completion due to the fact that the reactants and the product(s) reach a state of equilibrium. The following reactions, which may be used in the synthesis of methanol, are examples of equilibrium limited chemical reactions:



Please replace the paragraph beginning at page 9, line 7, with the following rewritten paragraph:

- - The term "primary reactant" refers to one of the reactants in a chemical reaction. The primary reactant may or may not be present at the highest concentration of the reactants in the reactant composition. An example of a primary reactant is CO in the above-indicated methanol synthesis reaction represented by Equation ~~(2)~~ (3).